**Additional Practice Problems - Unit 5 Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**\*\*\* Complete the following questions on a piece of notebook paper. You will need to show the completed problems to me before I allow you to retake the quiz!**

**Quiz #1 – Properties & Naming Ionic & Covalent Compounds**

1. Place an “X” in the correct box for each property in the table below:

|  |  |  |
| --- | --- | --- |
| **Property** | **Ionic** | **Covalent** |
| High melting point |  |  |
| Conducts when dissolved |  |  |
| Named using prefixes |  |  |
| Made of neutral atoms |  |  |
| Single molecules |  |  |
| Metal + nonmetal |  |  |
| Electrons transferred |  |  |
| Solids, liquids, & gases |  |  |

1. Name the following compounds:
   1. CaO
   2. IF4
   3. B6Si
   4. Cu(NO3)2
2. Write the formula for the following compounds:
   1. Phosphorous tribromide
   2. Iron (II) cyanide
   3. Cesium nitride
   4. Tetraphosphorous pentoxide

**Quiz #2 – Draw 1 Lewis Structure**

1. Draw the Lewis structure for H2Se.
2. Draw the Lewis structure for PBr3.

**Quiz #3 – Lewis Structures – All Types** Draw the Lewis Structure for each of the following covalent compounds:  
 1. PBr3 2. SiS2 3. SO3 4. NO31-

**Quiz #4 – Shapes & Polarity**

Draw the Lewis structure, determine the shape, and decide whether each of the following   
 molecules is polar or nonpolar:  
 1. PI3  
 2. CS2  
 3. NH41+  
 4. HBr